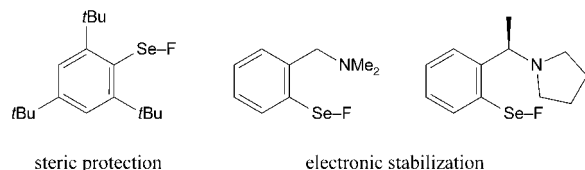


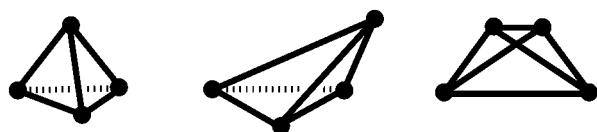
**Selenenyl fluorides ArSe–F**, generated by the reaction of Ar<sub>2</sub>Se<sub>2</sub> or ArSe–SiMe<sub>3</sub> with XeF<sub>2</sub>, are detected for the first time by <sup>77</sup>Se and <sup>19</sup>F NMR measurements (see picture). Other types of ArSe–F are unstable and disproportionate to ArSeF<sub>3</sub> and Ar<sub>2</sub>Se<sub>2</sub> (Ar = Ph, *o*-PrC<sub>6</sub>H<sub>4</sub>, *o*-MeOCH<sub>2</sub>C<sub>6</sub>H<sub>4</sub>, *o*-(S)-Et-(MeO)CHC<sub>6</sub>H<sub>4</sub>, 2,6-(MeOCH<sub>2</sub>)<sub>2</sub>C<sub>6</sub>H<sub>3</sub>).



H. Poleschner,\* K. Seppelt . . 6565–6574

First Detection of a Selenenyl Fluoride ArSe–F by NMR Spectroscopy: The Nature of Ar<sub>2</sub>Se<sub>2</sub>/XeF<sub>2</sub> and ArSe–SiMe<sub>3</sub>/XeF<sub>2</sub> Reagents

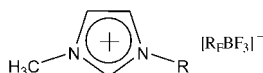
**The accepted mathematical conditions** for a degree of chirality of a geometrical object such as a tetrahedron lead to an infinite variety of such indices. Moreover, to every chiral geometric tetrahedron (see picture), no matter how close to an achiral or degenerate limit, there is some legitimate functional form for degree of chirality that makes this the “most chiral” tetrahedron.



A. Rassat,\* P. W. Fowler . . . 6575–6580

Is There a “Most Chiral Tetrahedron”?

**A series of new hydrophobic ionic liquids**, consisting of [R<sub>F</sub>BF<sub>3</sub>]<sup>−</sup> (R<sub>F</sub> = CF<sub>3</sub>, C<sub>2</sub>F<sub>5</sub>, nC<sub>3</sub>F<sub>7</sub>, nC<sub>4</sub>F<sub>9</sub>) anion with 1-alkyl(alkyl ether)-3-methylimidazolium ([C<sub>m</sub>mim]<sup>+</sup> or [C<sub>m</sub>O<sub>n</sub>mim]<sup>+</sup> (where C<sub>m</sub> is 1-alkyl, C<sub>m</sub> = nC<sub>m</sub>H<sub>2m+1</sub>, m = 1–4 and 6; C<sub>m</sub>O<sub>n</sub> is 1-alkyl ether, C<sub>2</sub>O<sub>1</sub> = CH<sub>3</sub>OCH<sub>2</sub>, C<sub>3</sub>O<sub>1</sub> = CH<sub>3</sub>OCH<sub>2</sub>CH<sub>2</sub>, and C<sub>5</sub>O<sub>2</sub> = CH<sub>3</sub>(OCH<sub>2</sub>CH<sub>2</sub>)<sub>2</sub>) cation; see structure), were synthesized and characterized. The key features for these new ionic liquids are their low melting points (−42 to 35°C) or extremely low glass transition (between −87 and −117°C) without melting, and considerably low viscosities (26–77 cP at 25°C).



Z.-B. Zhou, H. Matsumoto,\*

K. Tatsumi . . . . . 6581–6591

Low-Melting, Low-Viscous, Hydrophobic Ionic Liquids: 1-Alkyl(Alkyl Ether)-3-methylimidazolium Perfluoroalkyltrifluoroborate

Supporting information on the WWW (see article for access details).

\* Author to whom correspondence should be addressed

Full Papers labeled with this symbol have been judged by two referees as being “very important papers”.

All the Tables of Contents from 1998 onwards may be found on the WWW under <http://www.chemeurj.org>

If not otherwise indicated in the articles, papers in issue number 23, 2004, were published online under <http://www.interscience.wiley.com/> on November 18, 2004.

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## CORRIGENDUM

Contents

Contrary to the published acknowledgement, Professor R. F. W. Bader in no way contributed to any of arguments presented in this paper; he has indicated that all of the criticisms raised by the authors were previously discussed and rebutted in *J. Chem. Phys. A* **1998**, 102, 7314.

A copy of an earlier version of the manuscript was sent to Professor Bader last spring, and the e-mail correspondence that followed made it very clear that he disagreed strongly with the conclusions. The authors nevertheless found the discussion useful, since it clarified the nature of the disagreement. Professor Bader's queries also prompted the authors to calculate the dissociation energy of the He@adam inclusion complex and to determine the structure of the transition state as discussed in the article. The authors apologize for omitting to ask for Professor Bader's permission to include his name among the acknowledgements.

A. Haaland, D. J. Shorokhov, N. V. Tverdova . . . . . 4416–4421

Topological Analysis of Electron Densities: Is the Presence of an Atomic Interaction Line in an Equilibrium Geometry a Sufficient Condition for the Existence of a Chemical Bond?

*Chem. Eur. J.* **2004**, 10

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